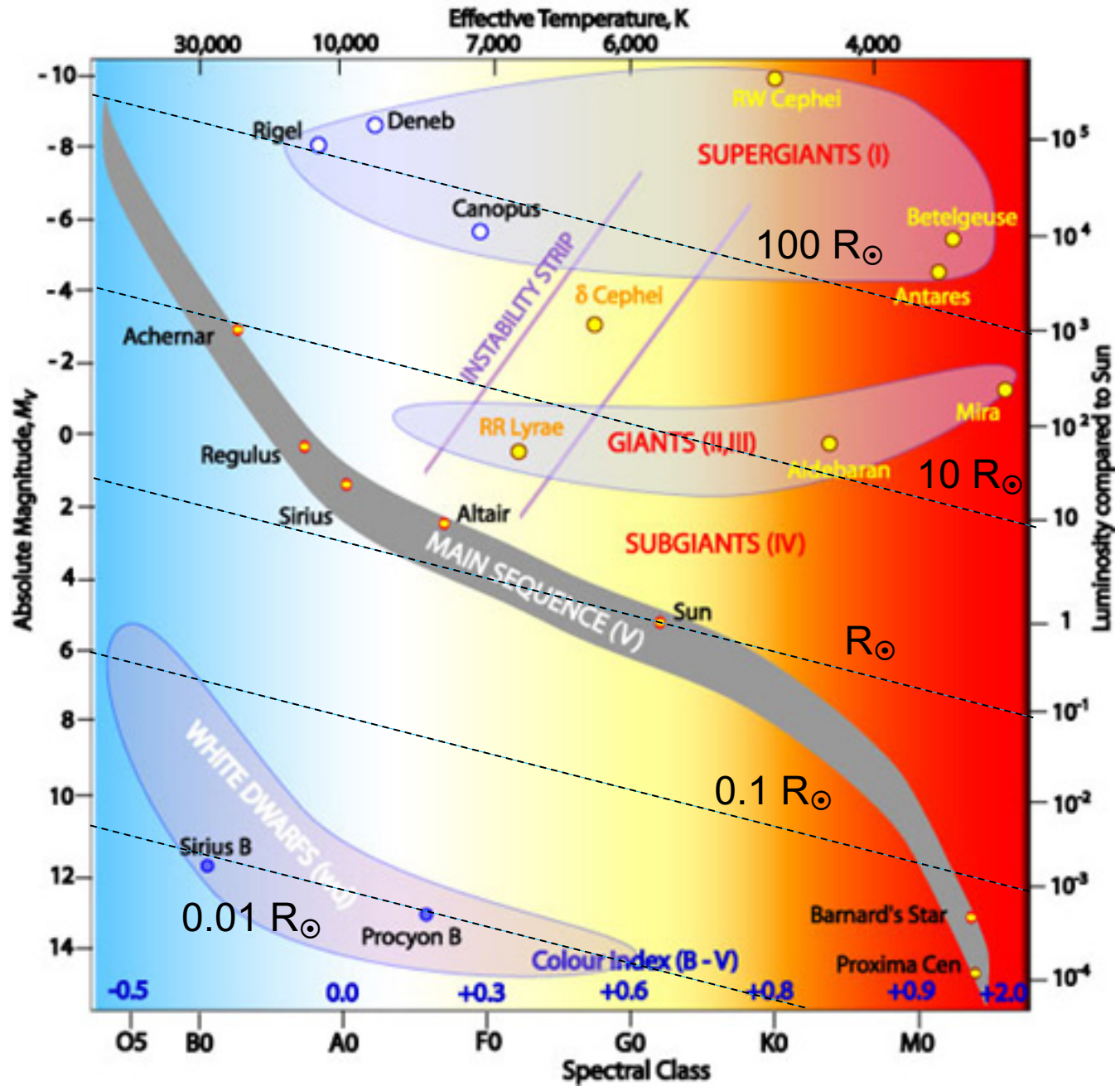
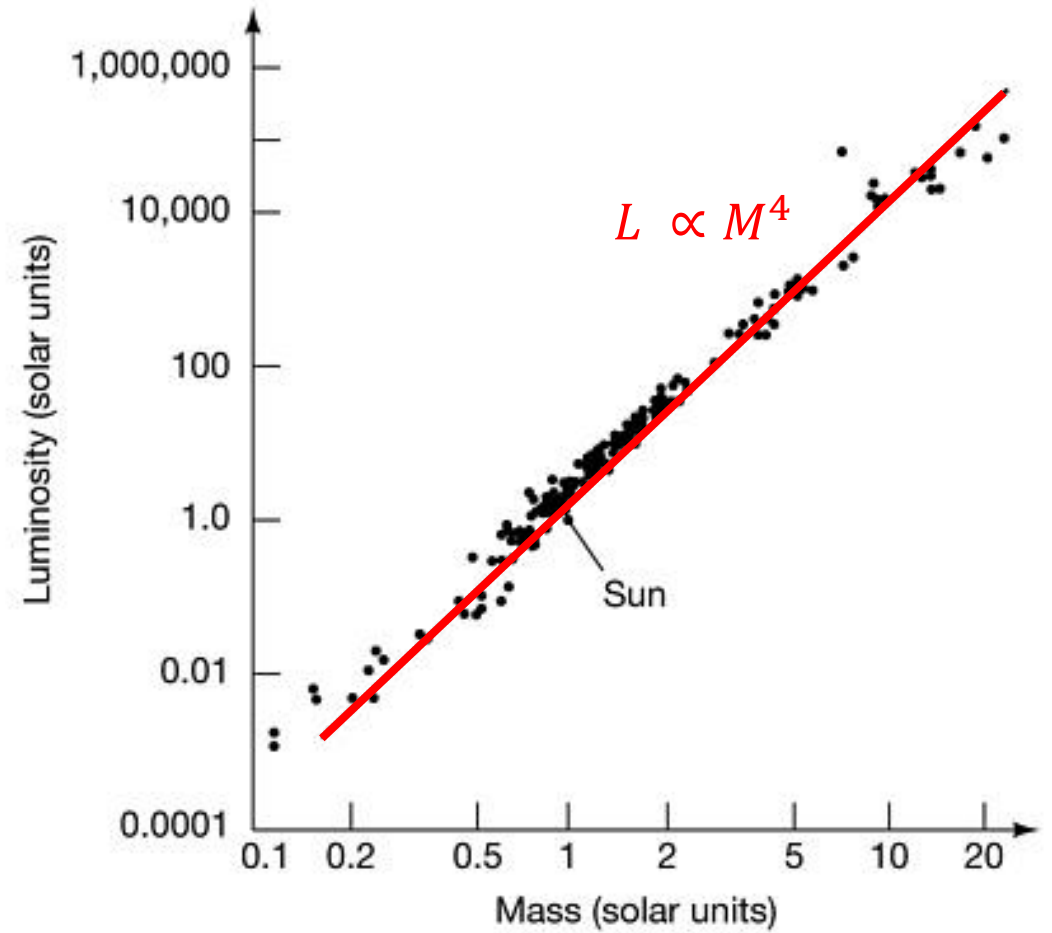
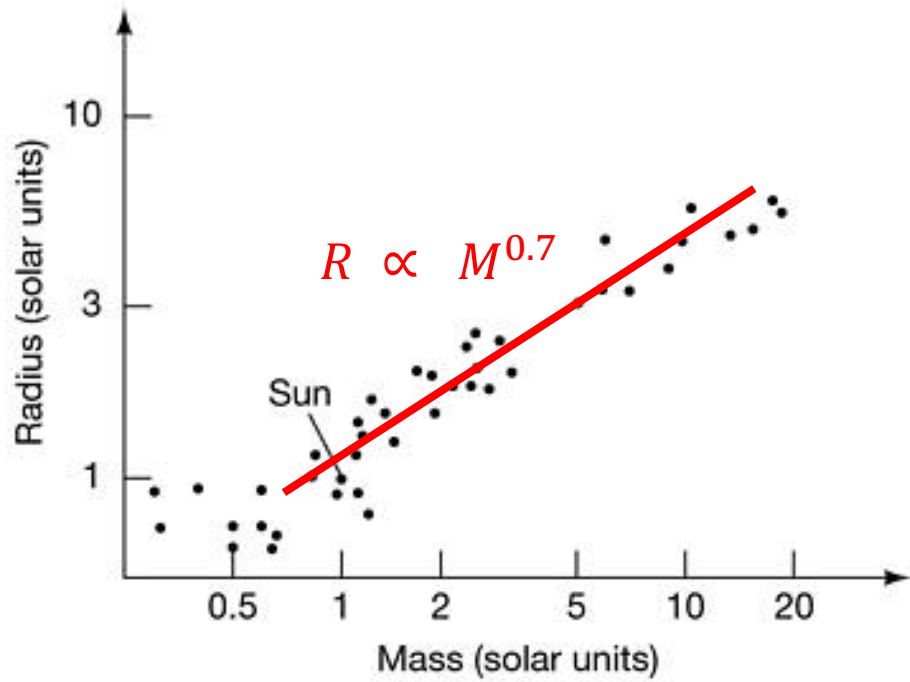


Hertzsprung-Russell Diagram





$$t_{MS} \sim \frac{M}{dM/dt} \propto \frac{M}{L} \propto M^{-3}$$

→ stars must evolve

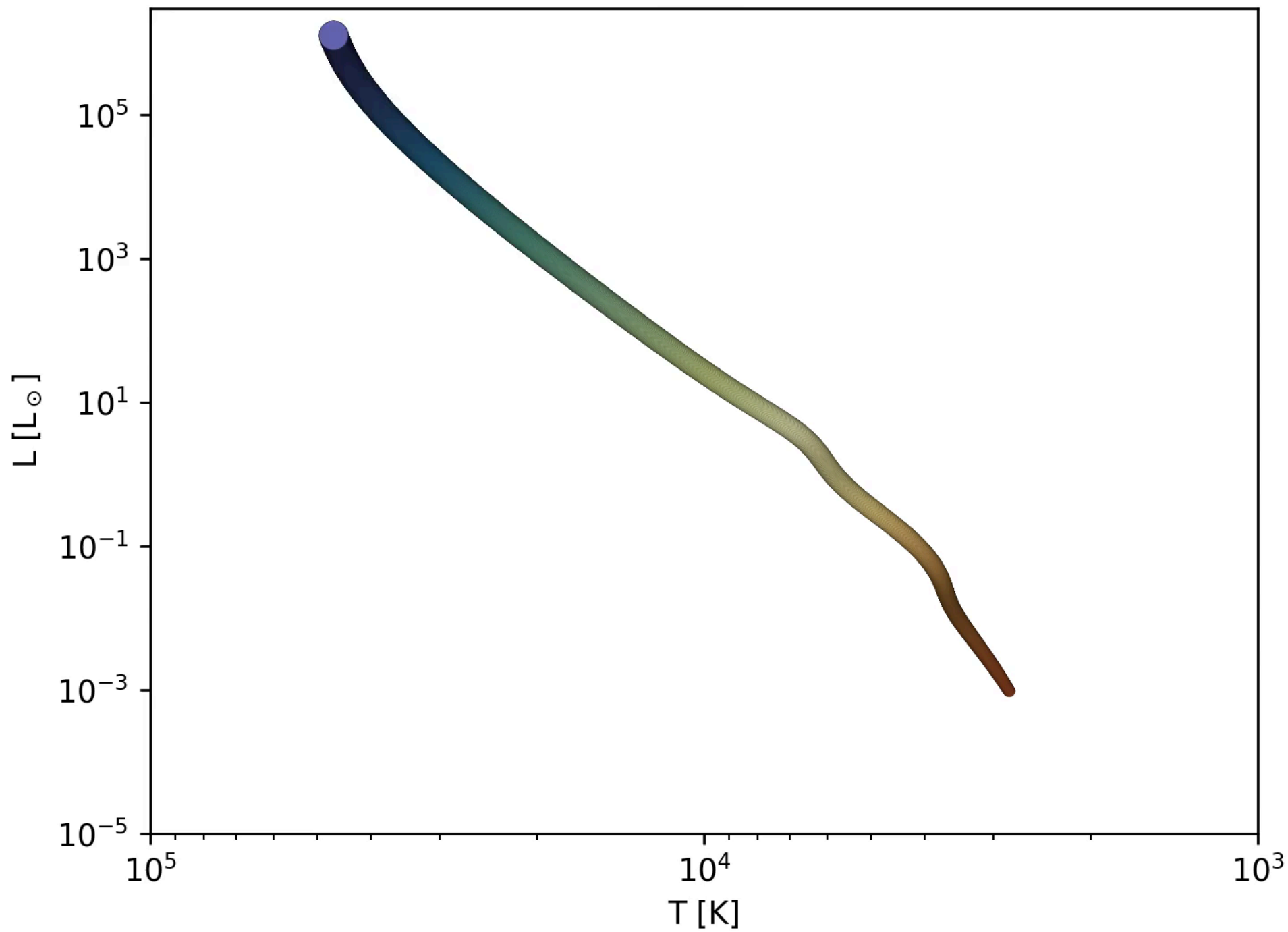
→ the most massive stars evolve fastest!

$$t_{MS} \approx 10^{10} \left(\frac{M}{M_{\odot}} \right)^{-3} \text{ years}$$

→ stars must evolve

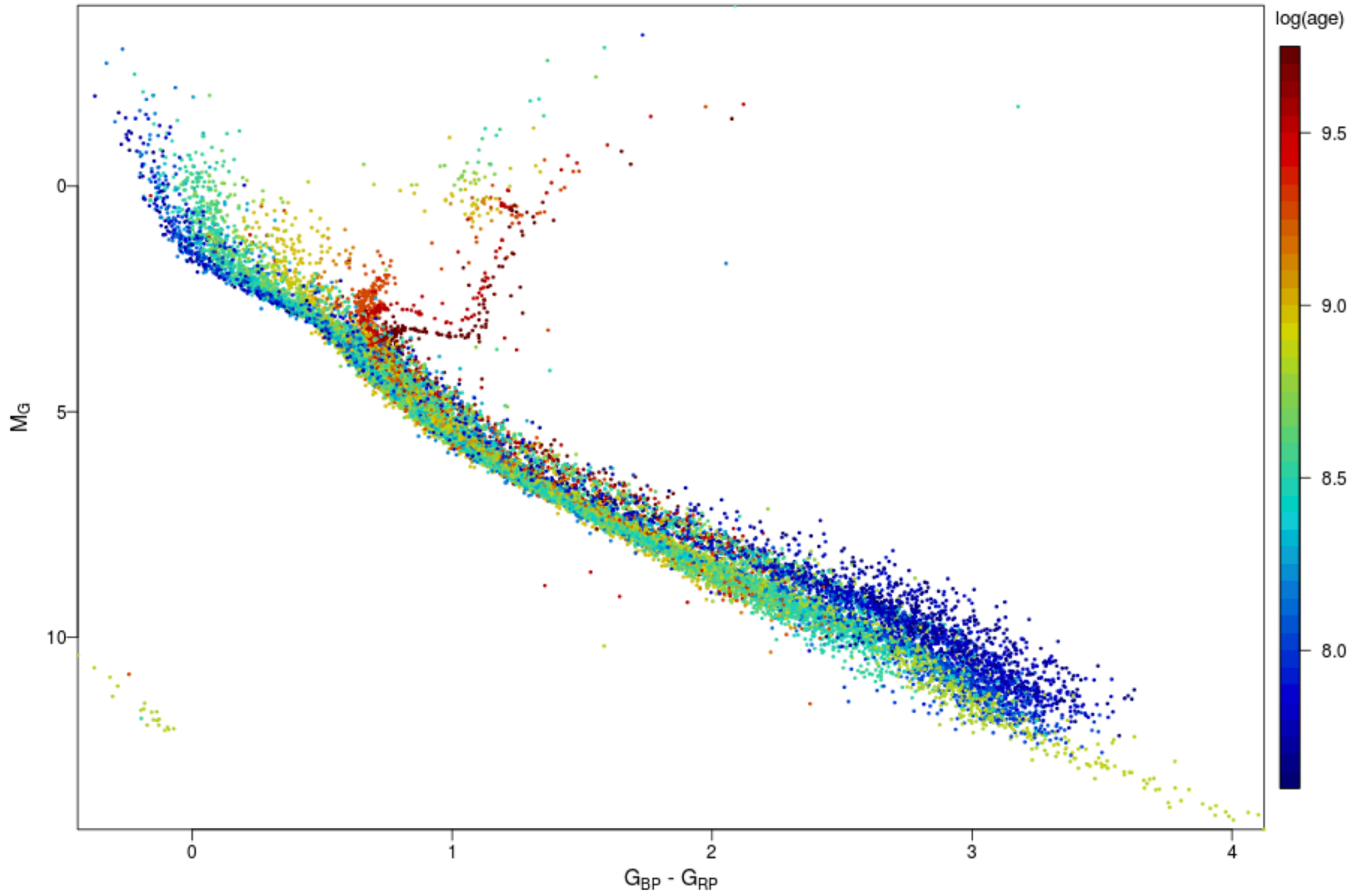
→ the most massive stars evolve fastest!

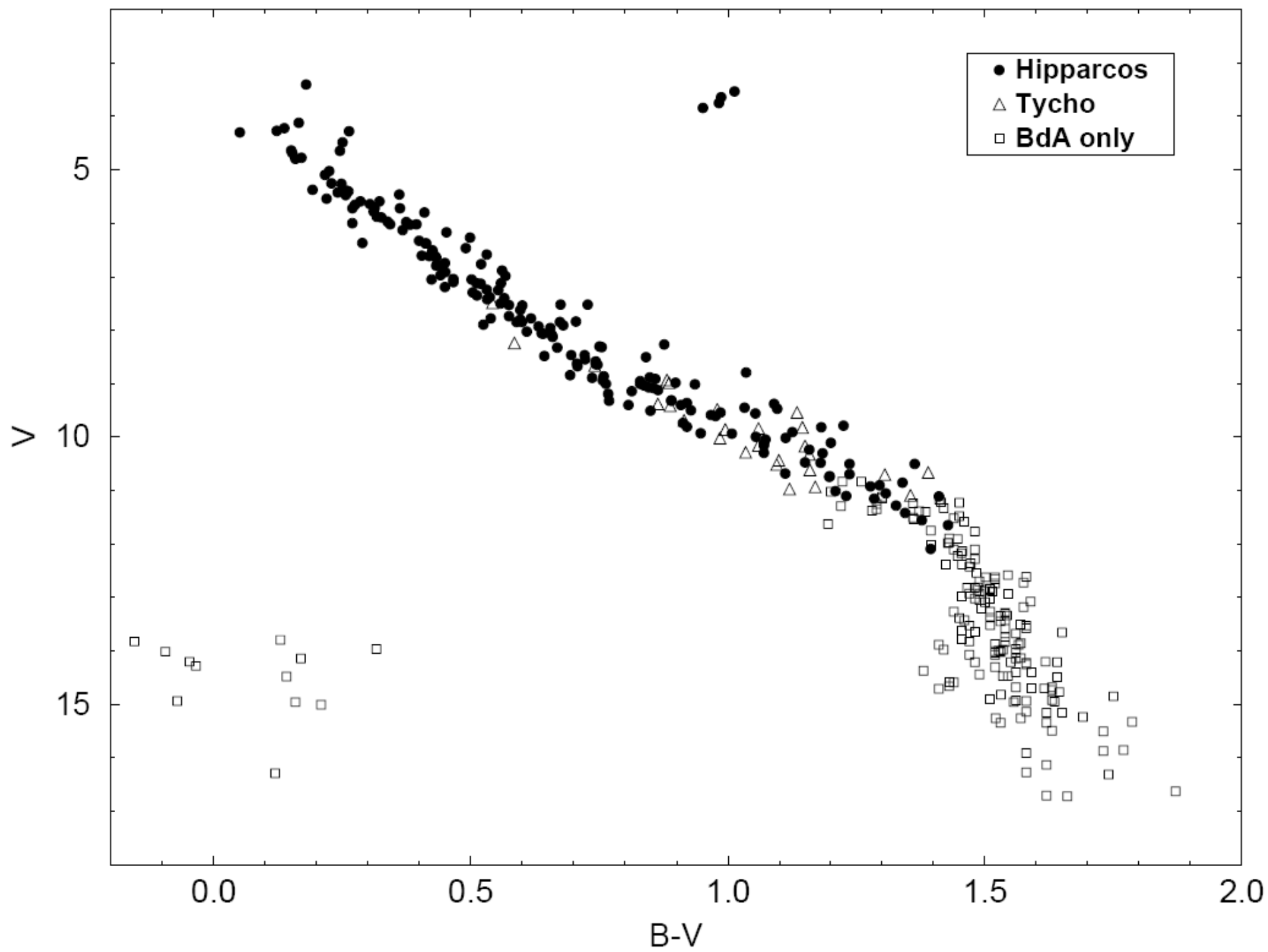
0.0 Myr



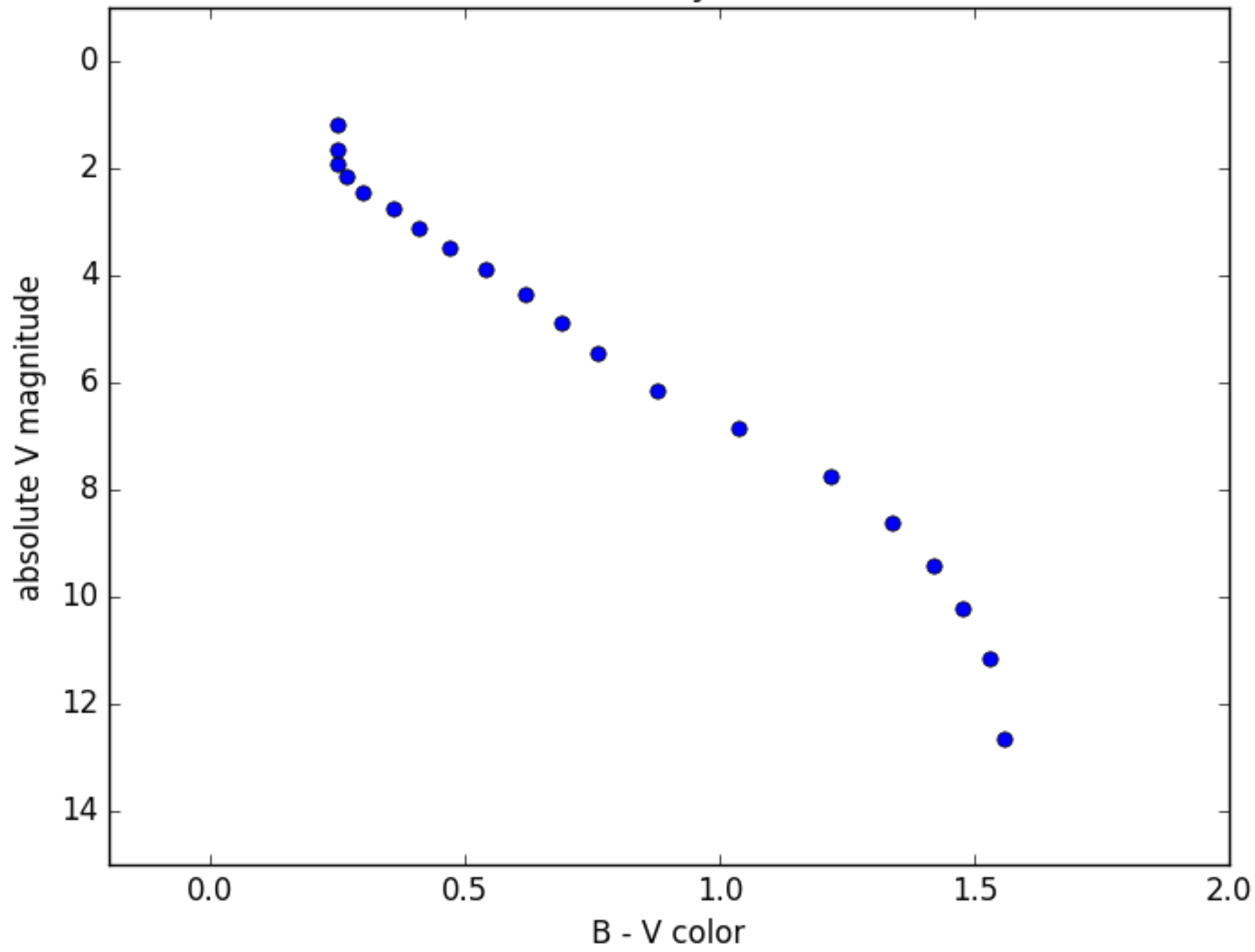
Gaia DR2

A&A 616, A10 (2018)

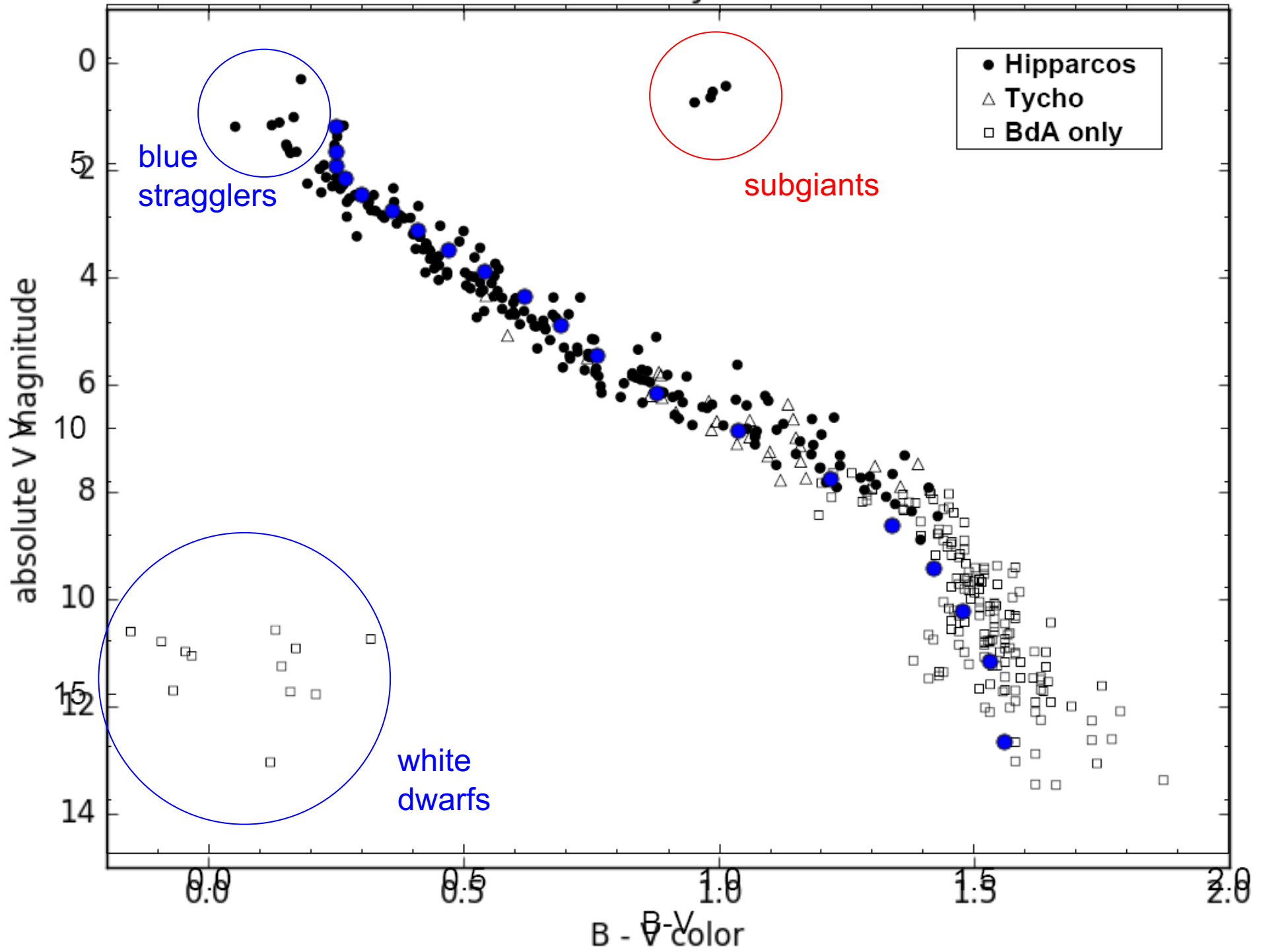


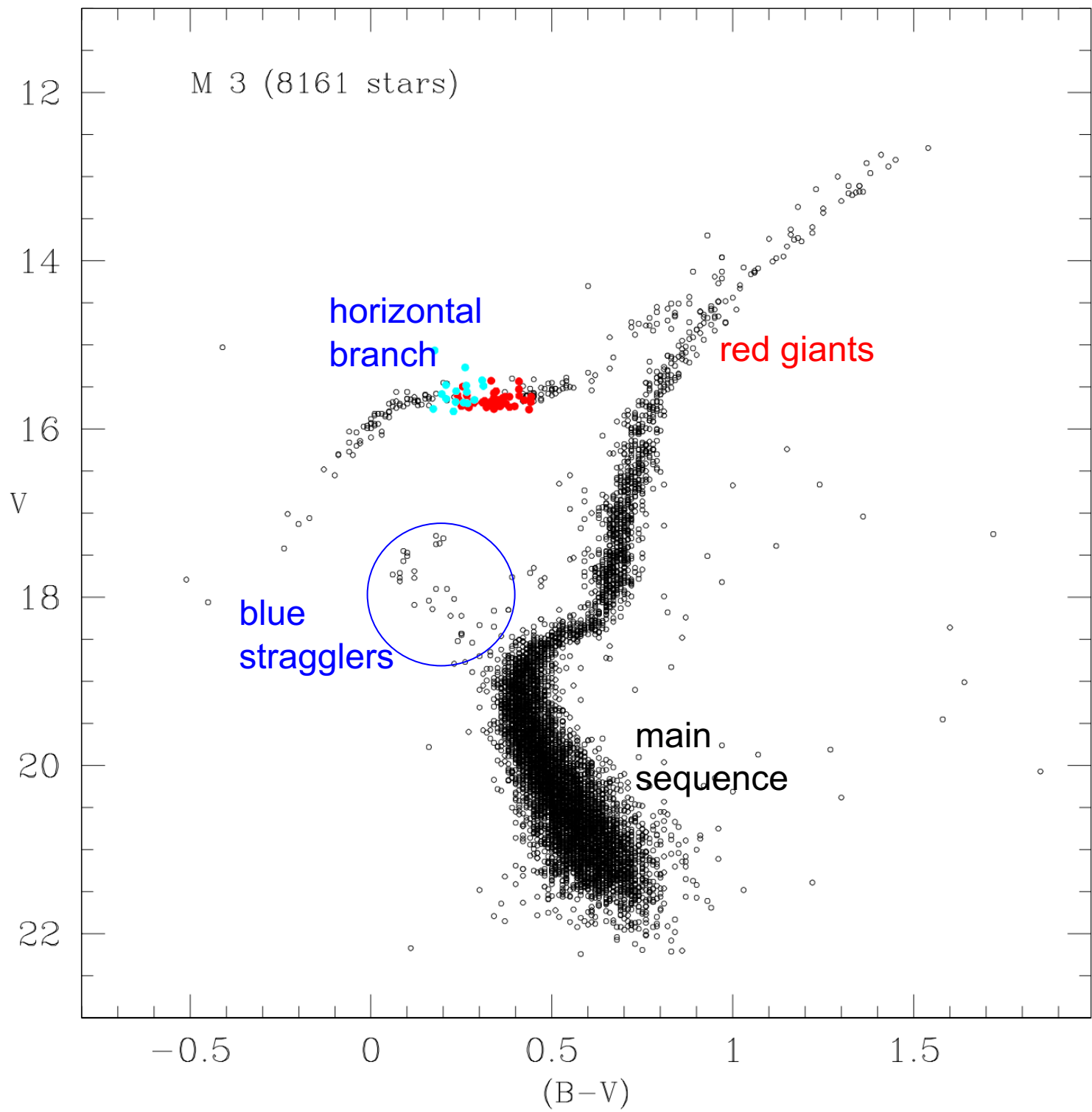


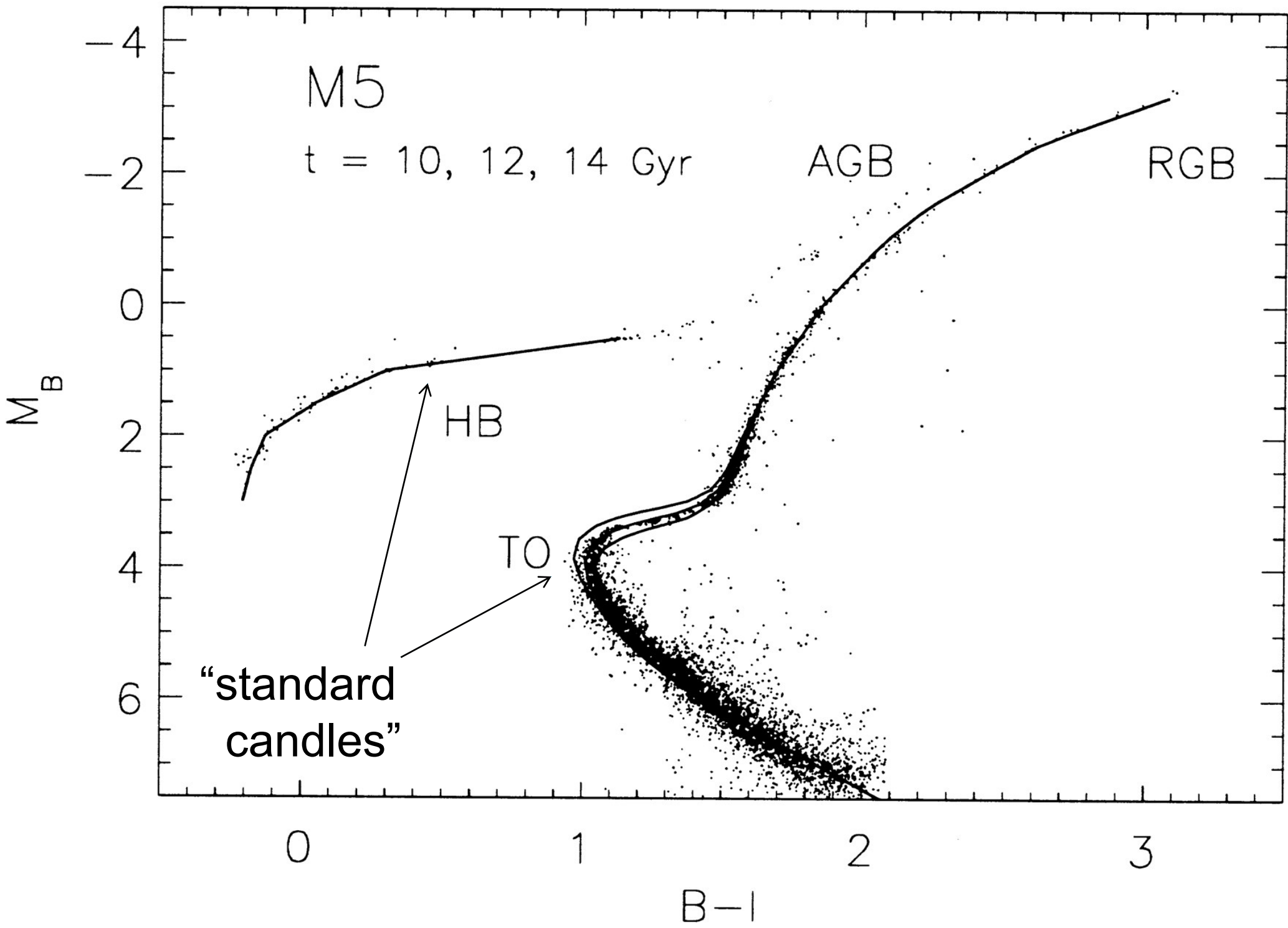
Siess et al. 1997 Hyades isochrone



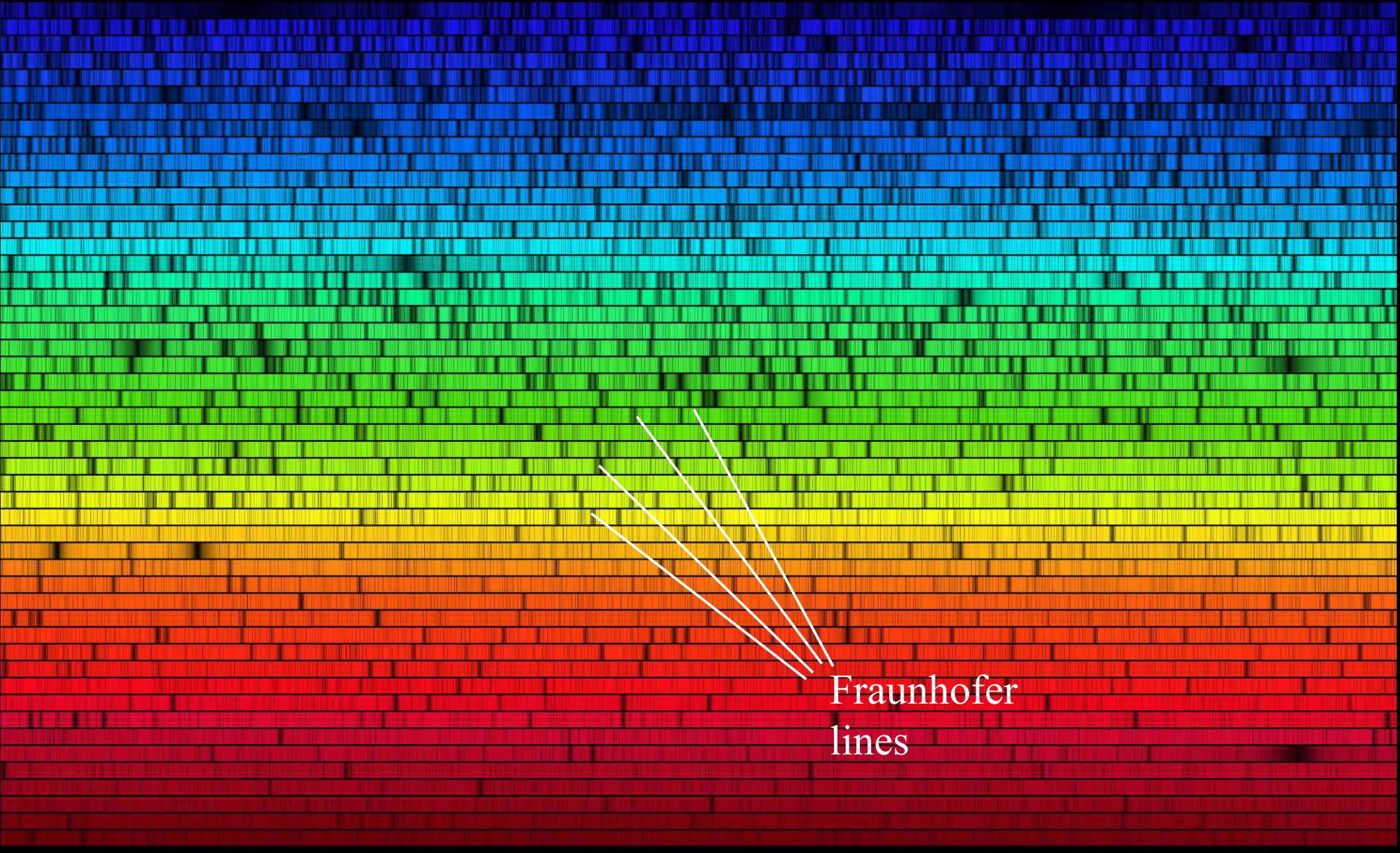
Siess et al. 1997 Hyades isochrone







Solar Spectrum



Continuous Spectrum

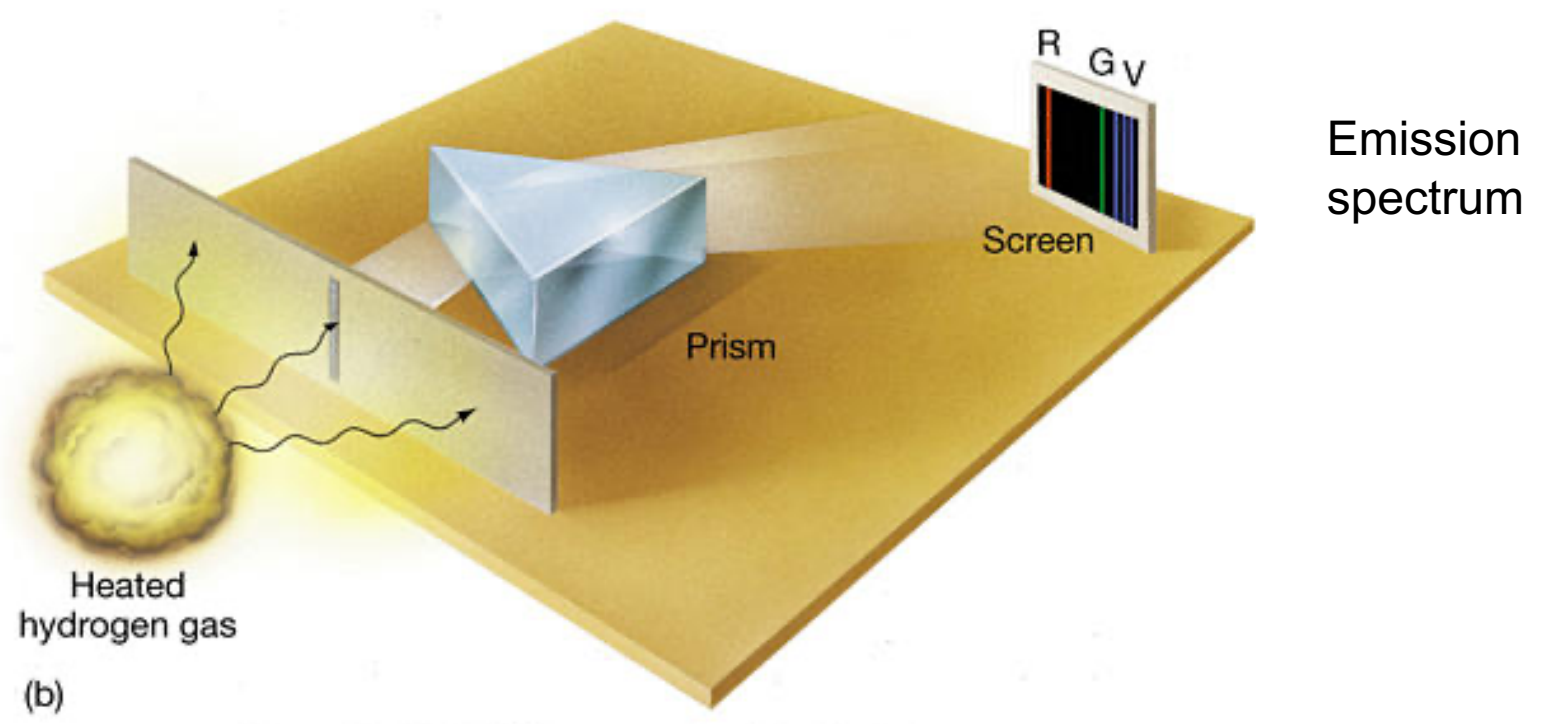
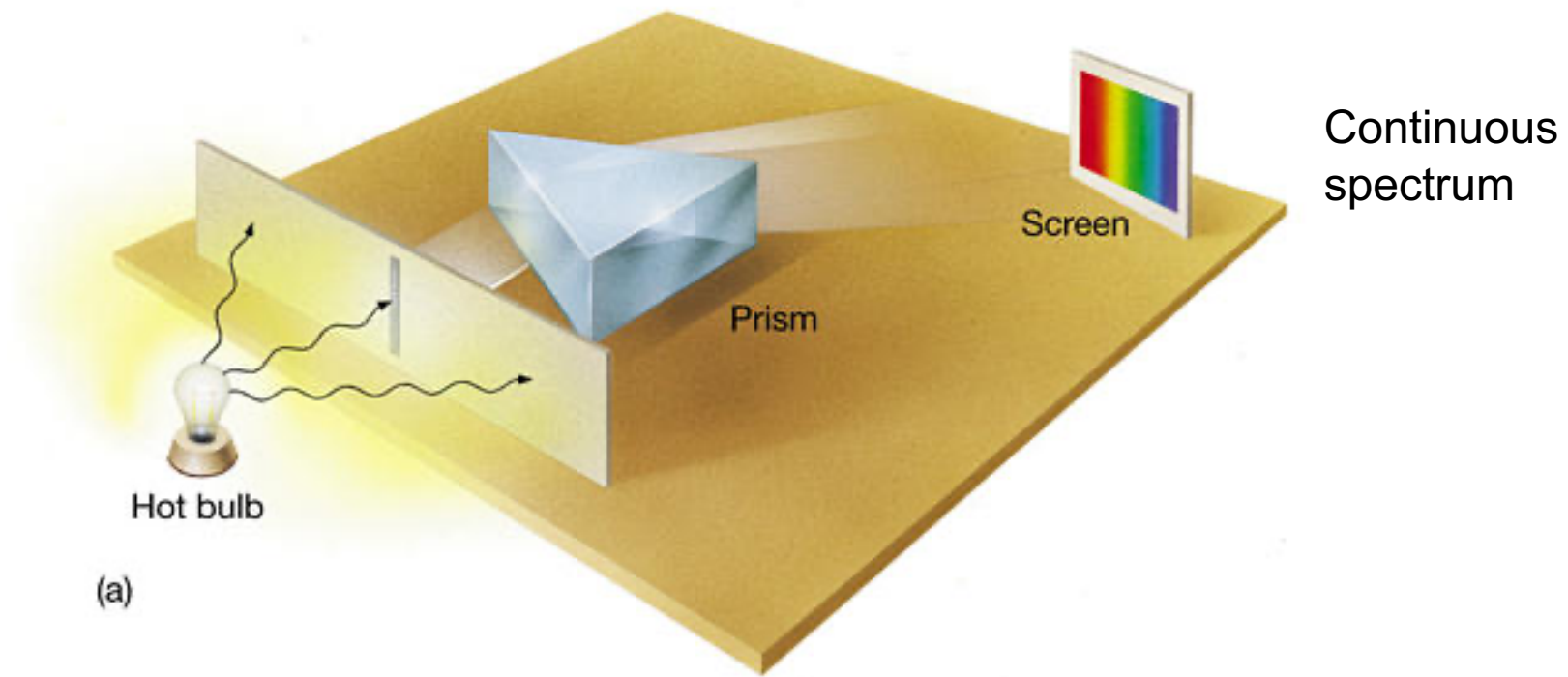


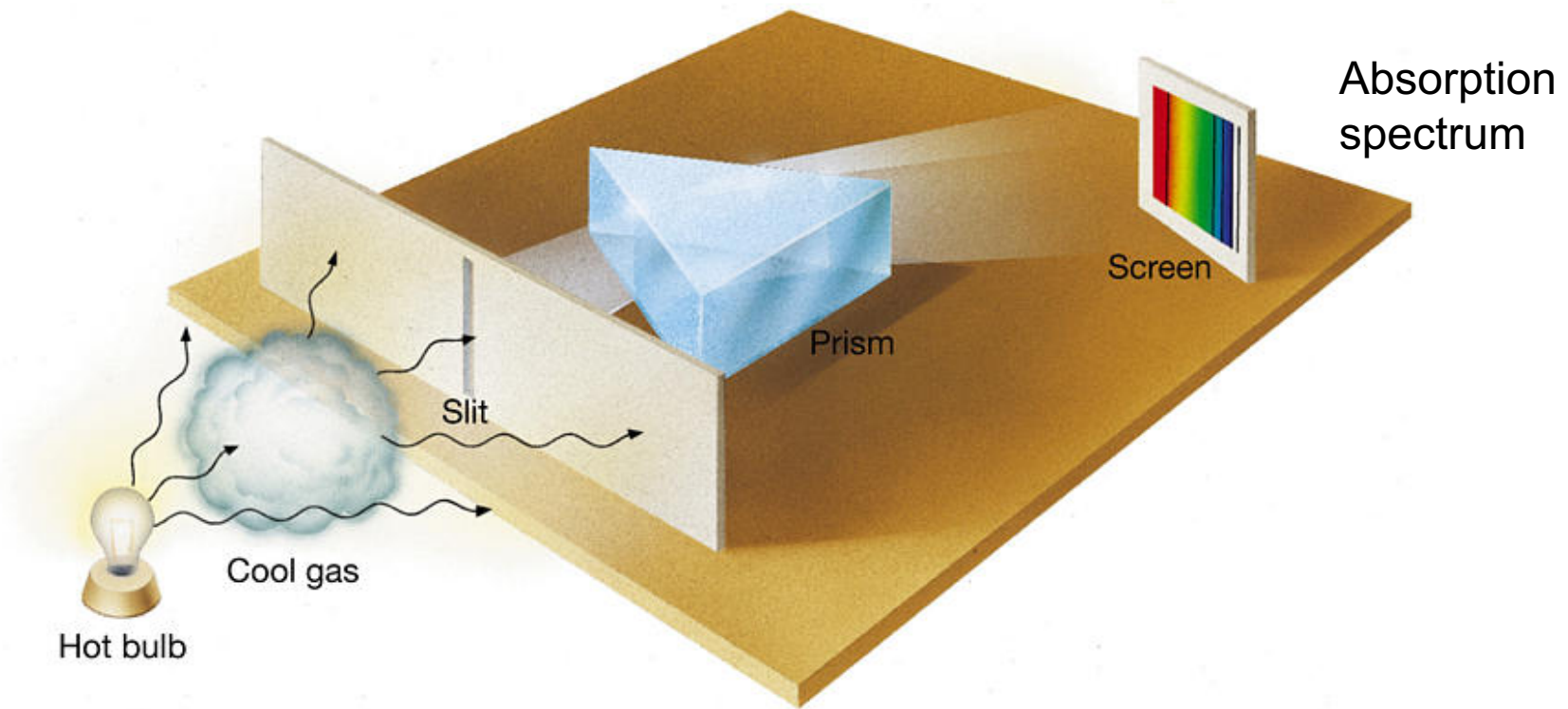
Emission Lines



Absorption Lines









Continuum



Hydrogen



Helium



Argon



Nitrogen



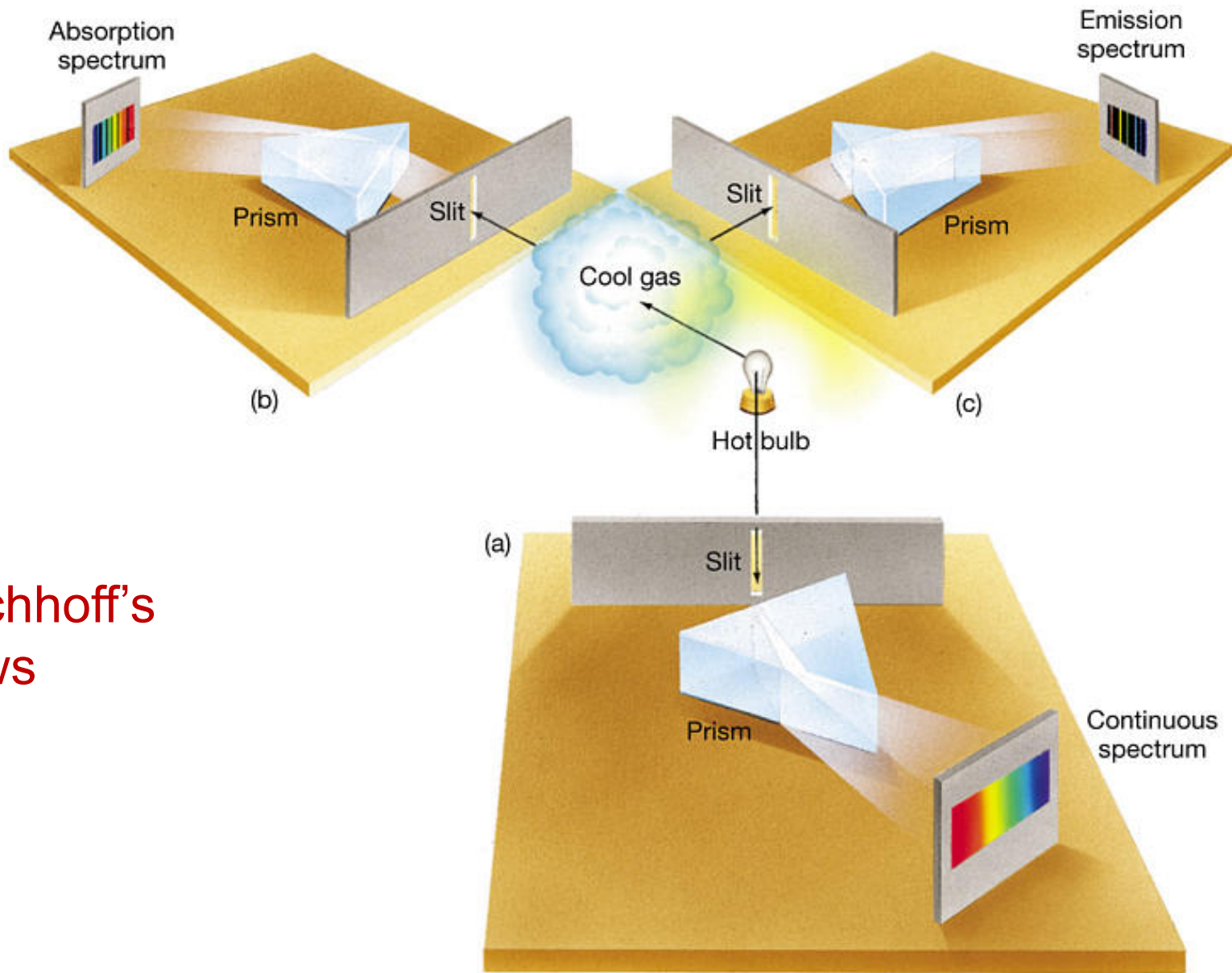
Neon



Mercury

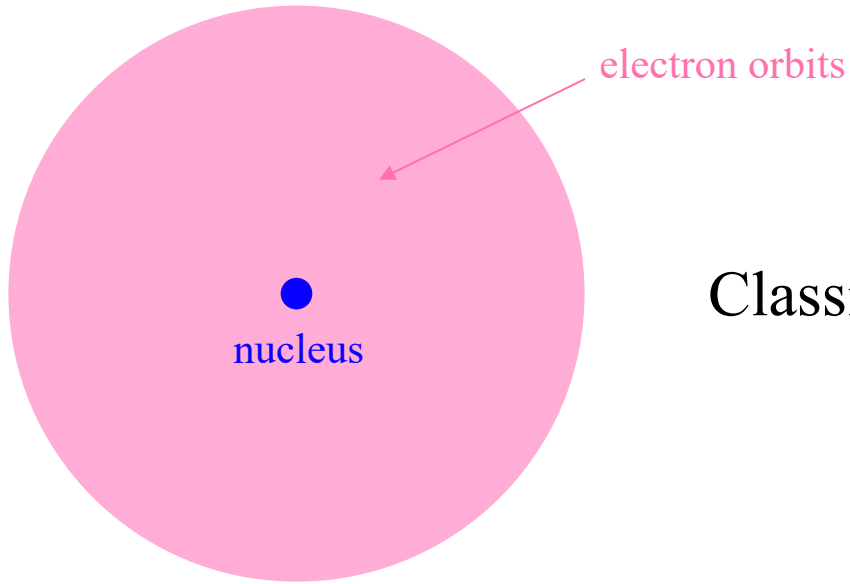


Carbon Monoxide

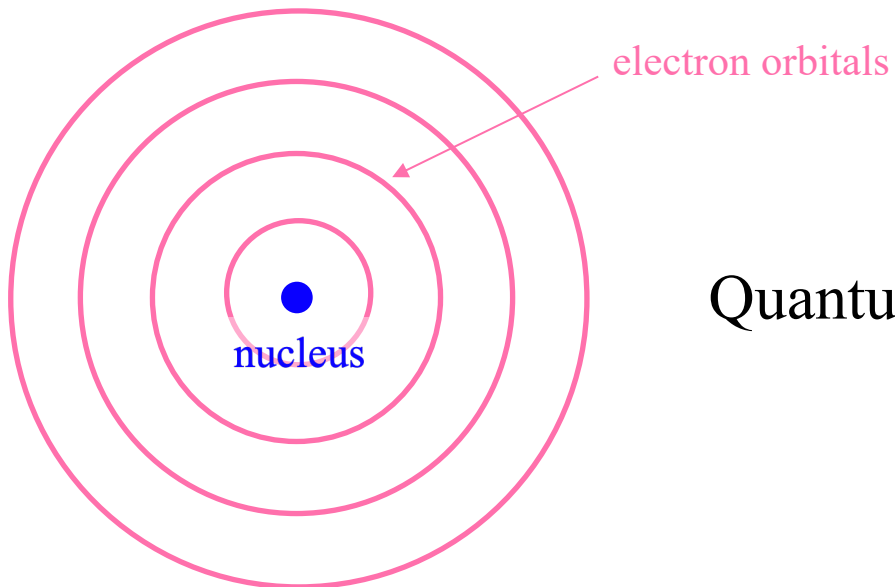


Kirchhoff's Laws

Quantization of Electron Energies

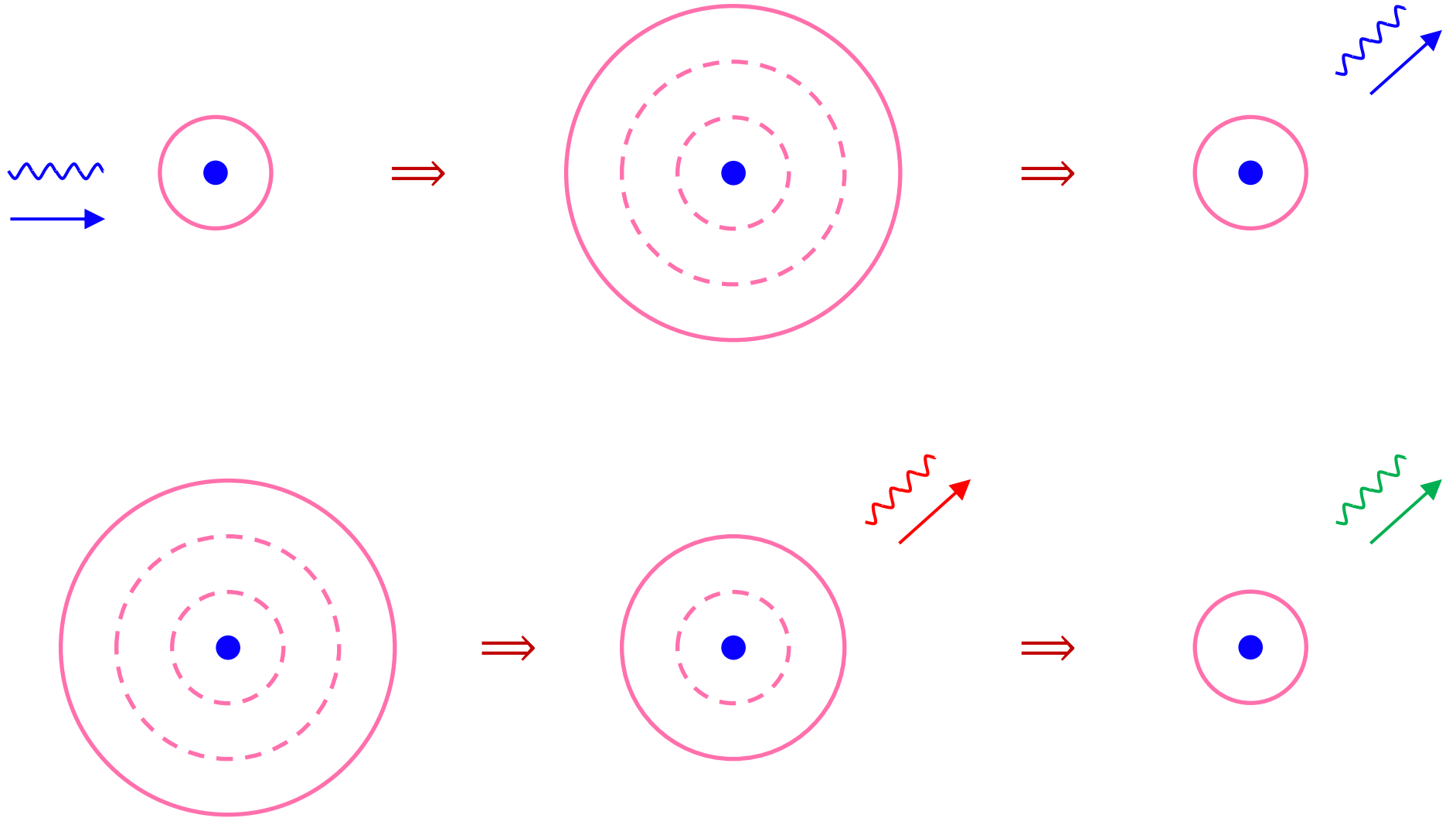


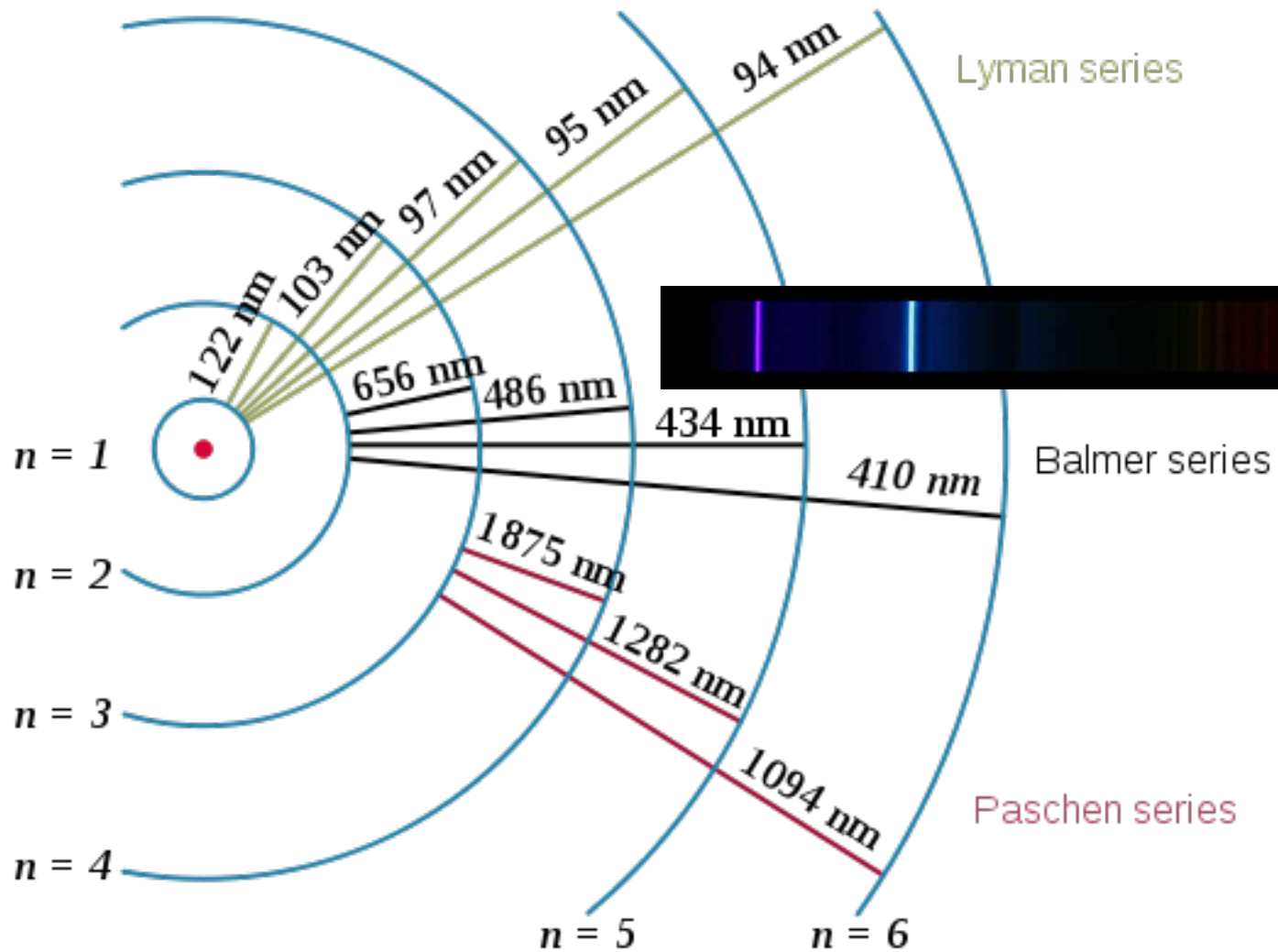
Classical picture: all electron orbits allowed



Quantum picture: only specific electron orbits allowed

Photon-Atom Interaction





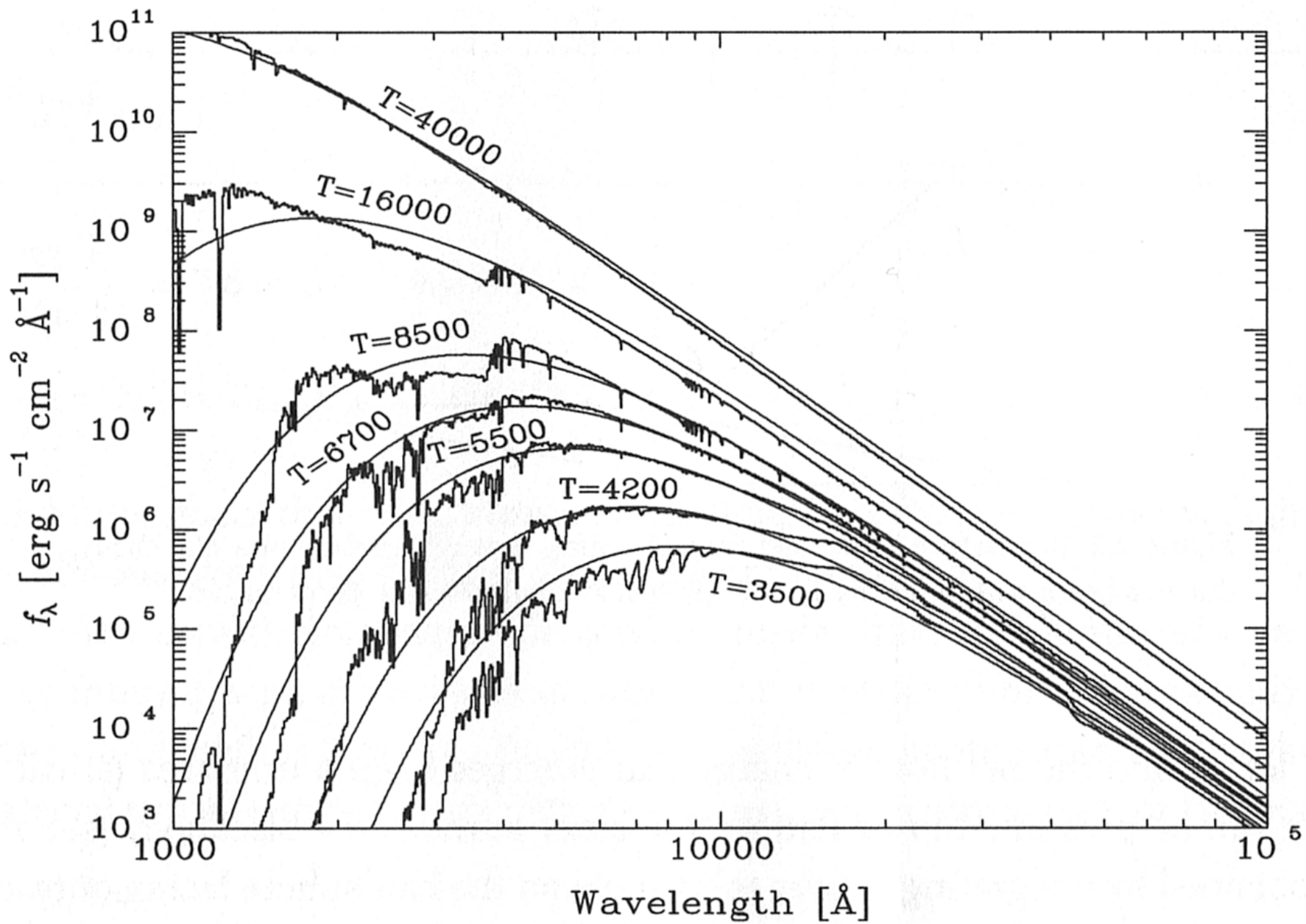
$$E_n = -13.6 \text{ eV} / n^2$$

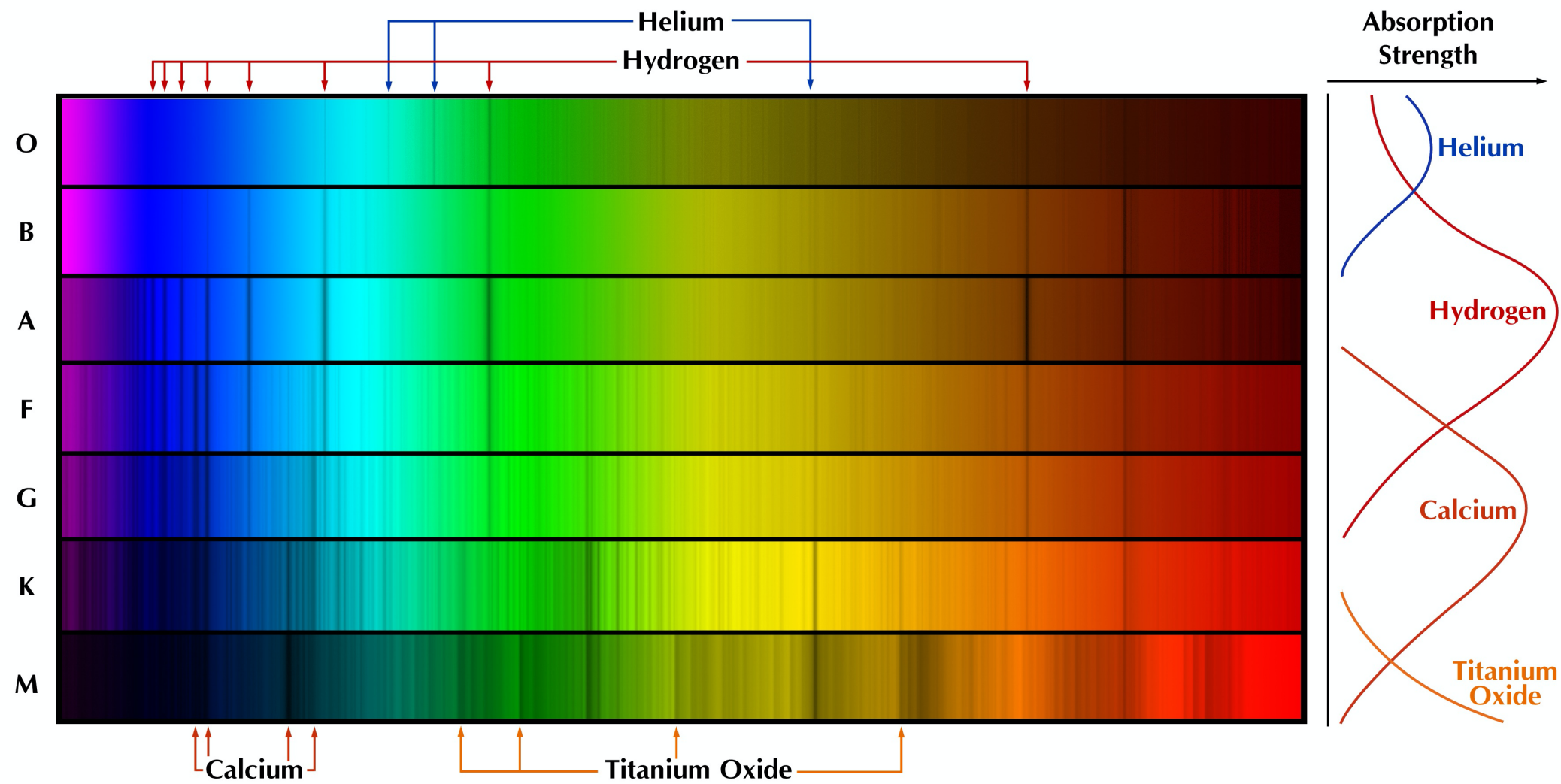
$$1 \text{ eV} = 1.6 \times 10^{-19} \text{ J}$$

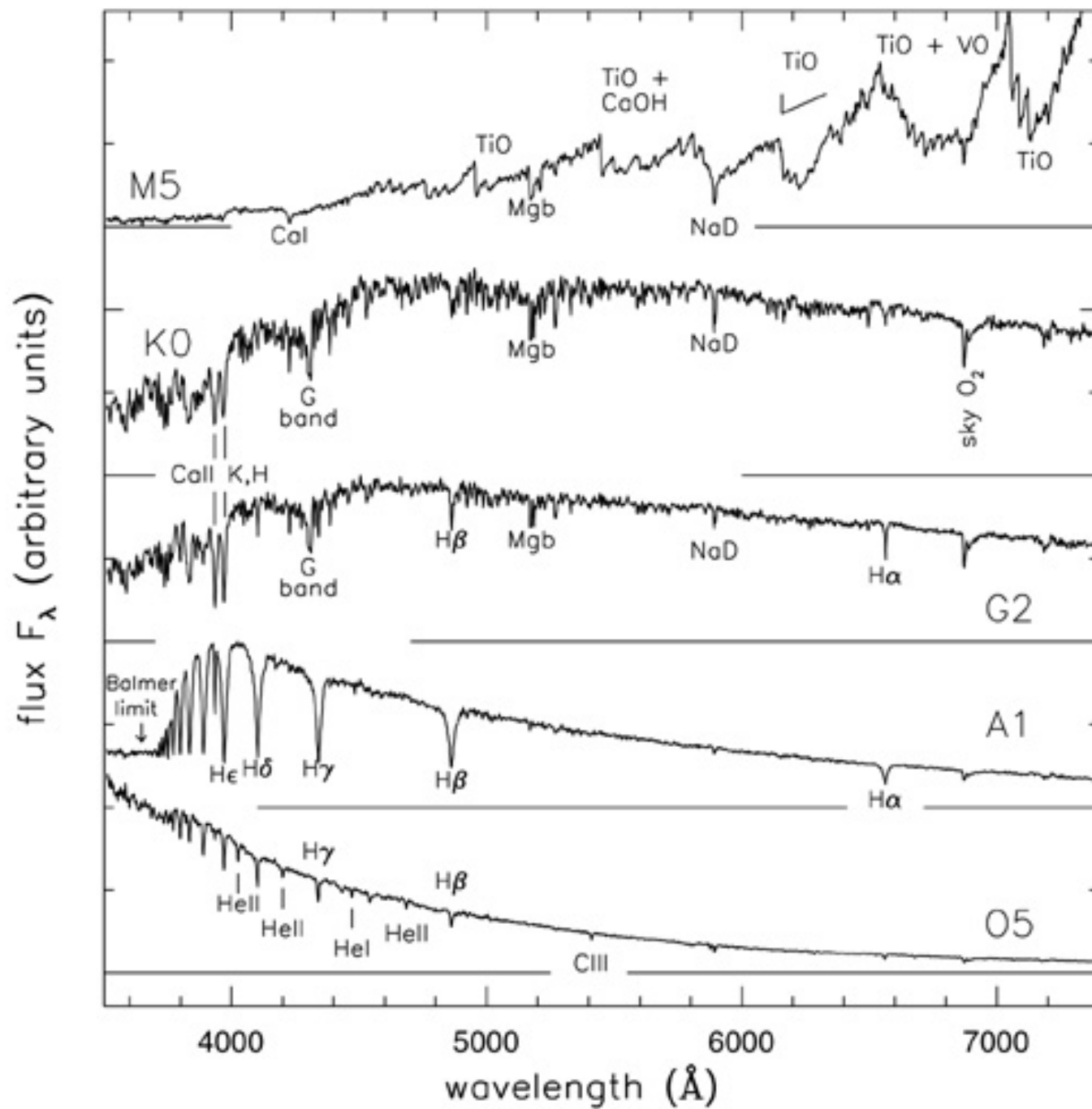
$$h\nu = E_n - E_m$$

Solar Composition

Element	% by Volume	% by Mass
Hydrogen	90.946	70.682
Helium	8.913	27.509
Oxygen	0.077	0.954
Carbon	0.033	0.303
Neon	0.011	0.170
Nitrogen	0.010	0.108
Magnesium	0.004	0.068
Iron	0.003	0.137
Silicon	0.003	0.069

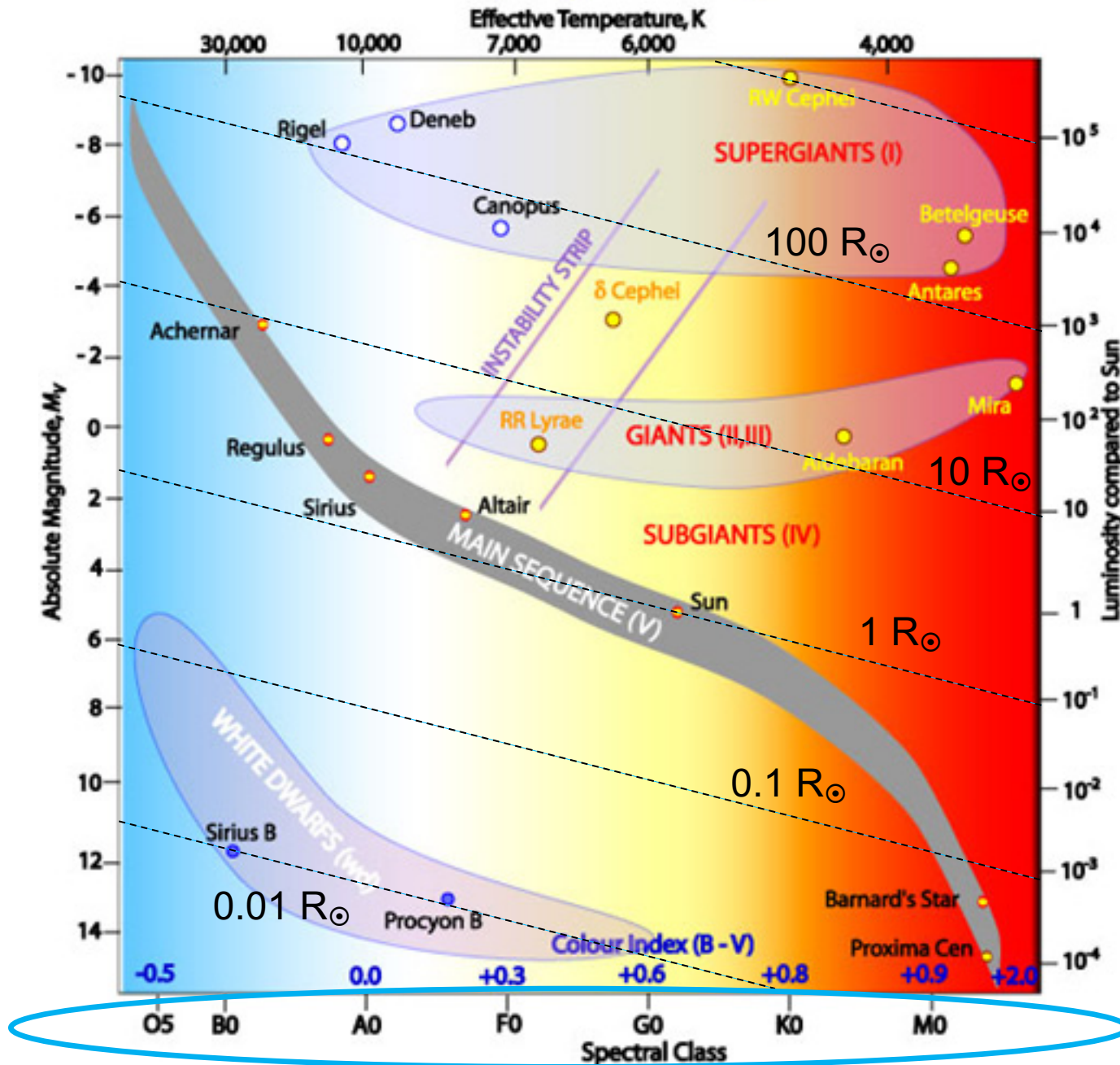






Spectral Class	Intrinsic Color	Temperature (K)	Prominent Absorption Lines
O	Blue	41,000	He ⁺ , O ⁺⁺ , N ⁺⁺ , Si ⁺⁺ , He, H
B	Blue	31,000	He, H, O ⁺ , C ⁺ , N ⁺ , Si ⁺
A	Blue-white	9,500	H(strongest), Ca ⁺ , Mg ⁺ , Fe ⁺
F	White	7,240	H(weaker), Ca ⁺ , ionized metals
G	Yellow-white	5,920	H(weaker), Ca ⁺ , ionized & neutral metal
K	Orange	5,300	Ca ⁺ (strongest), neutral metals strong, H(weak)
M	Red	3,850	Strong neutral atoms, TiO

Hertzsprung-Russell Diagram

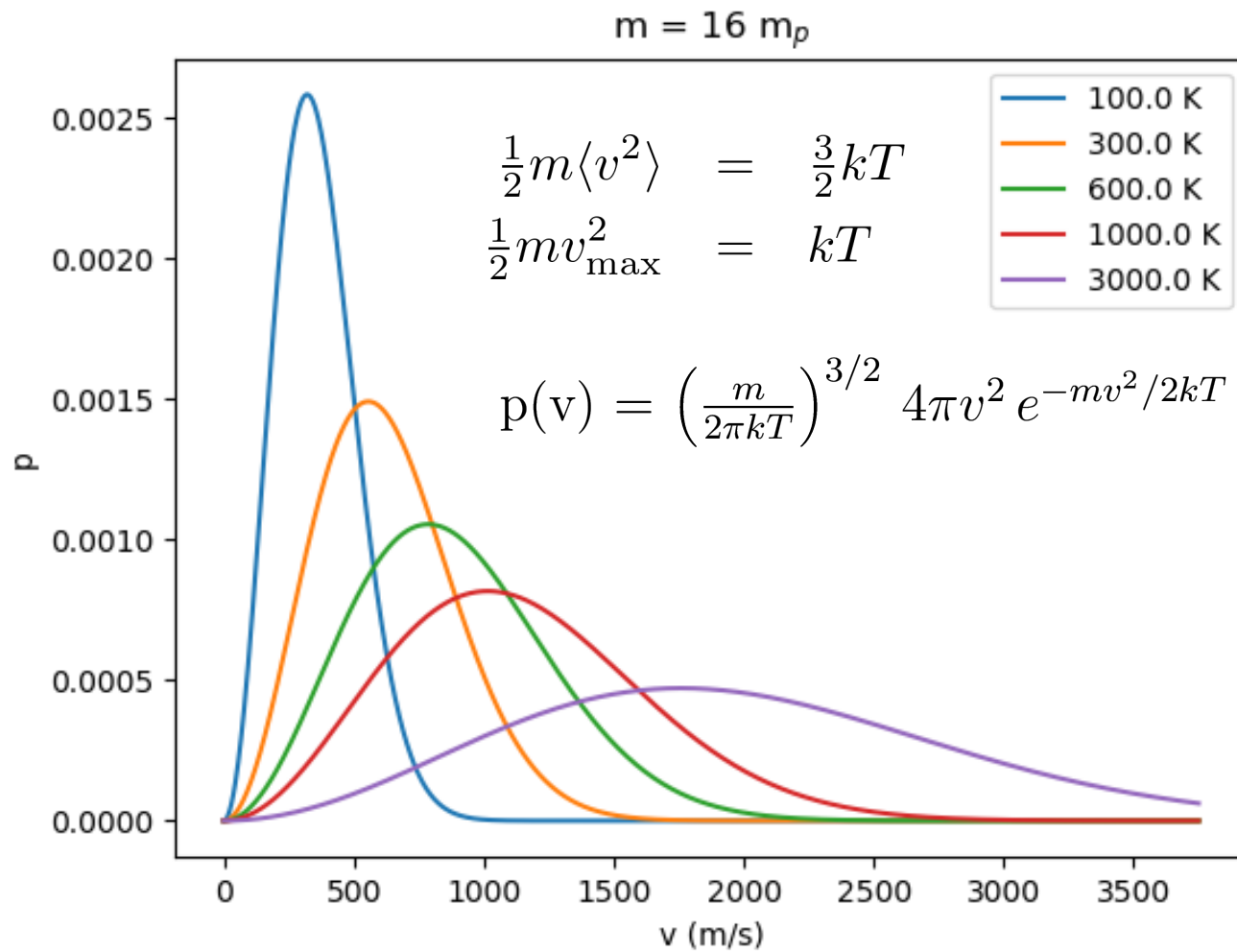


Maxwell-Boltzmann Distribution

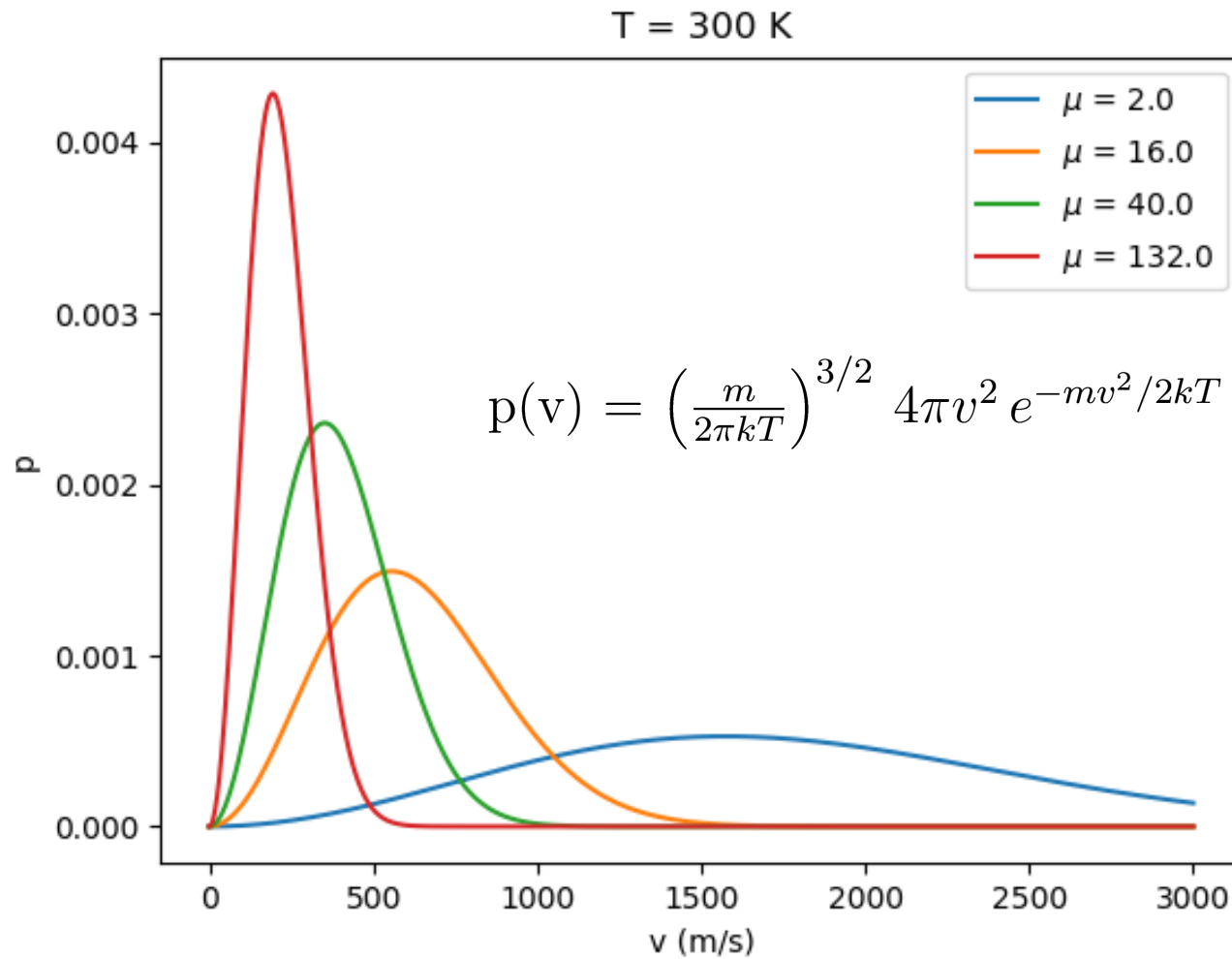
Most probable distribution of particle (mass m) velocities in thermal equilibrium at temperature T

$$p(v) = \left(\frac{m}{2\pi kT} \right)^{3/2} 4\pi v^2 e^{-mv^2/2kT}$$

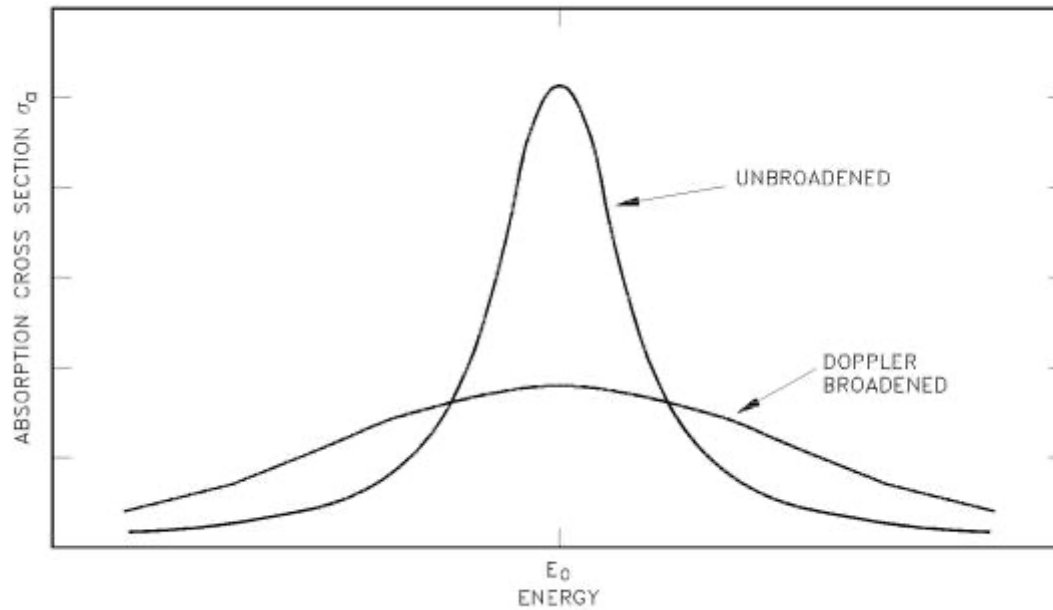
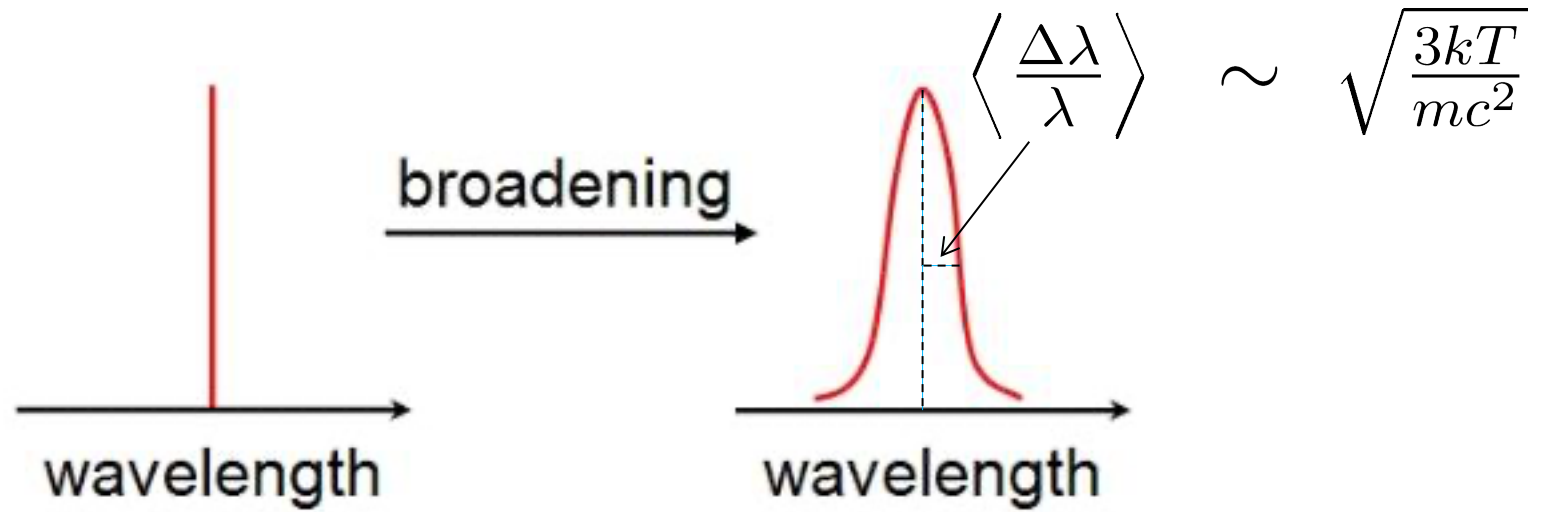
Maxwell-Boltzmann Distribution



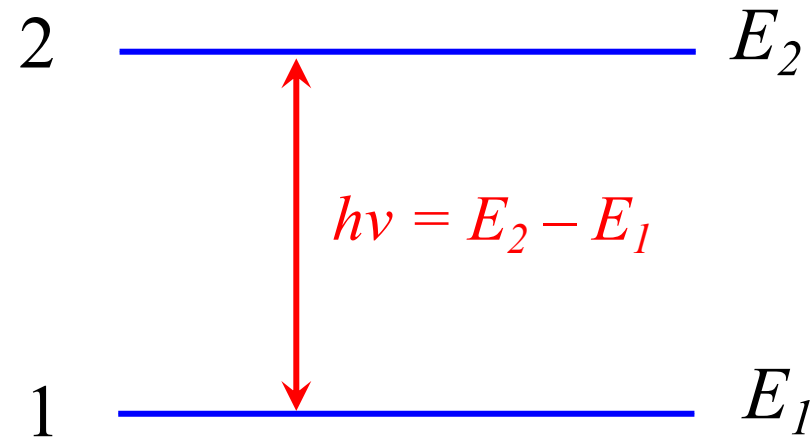
Maxwell-Boltzmann Distribution



Line Broadening



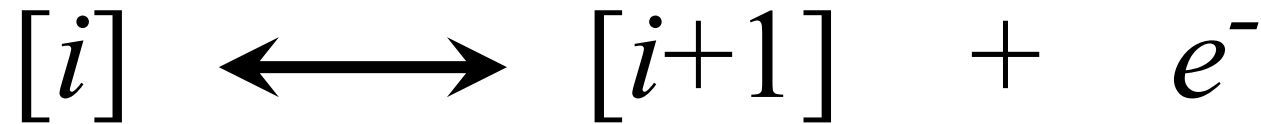
Boltzmann Equation



$$\frac{n_2}{n_1} = \frac{g_2}{g_1} e^{-(E_2 - E_1)/kT}$$

Hydrogen: $g_n = n^2$

Saha Equation

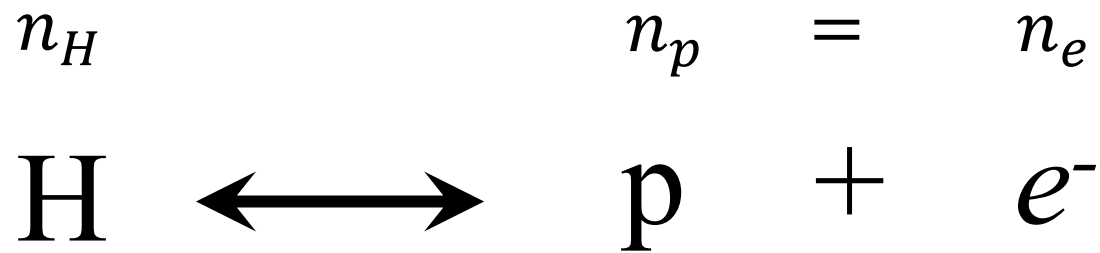


$$\frac{n_{i+1}n_e}{n_i} = \frac{2}{\Lambda^3} \frac{g_{i+1}}{g_i} e^{-\chi_i/kT}$$

excitation
potential

$$\Lambda = \sqrt{\frac{h^2}{2\pi m_e kT}}$$

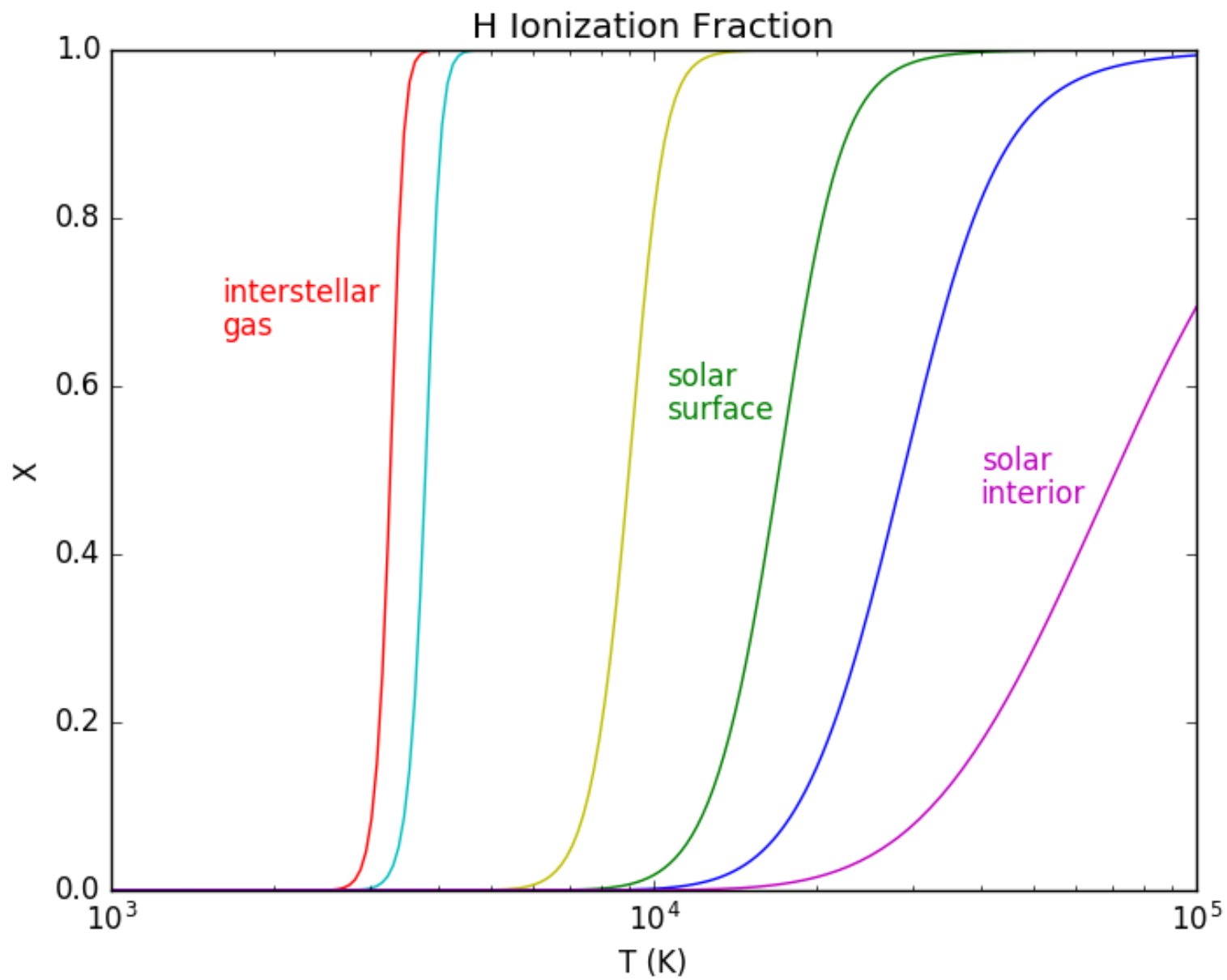
Hydrogen ionization

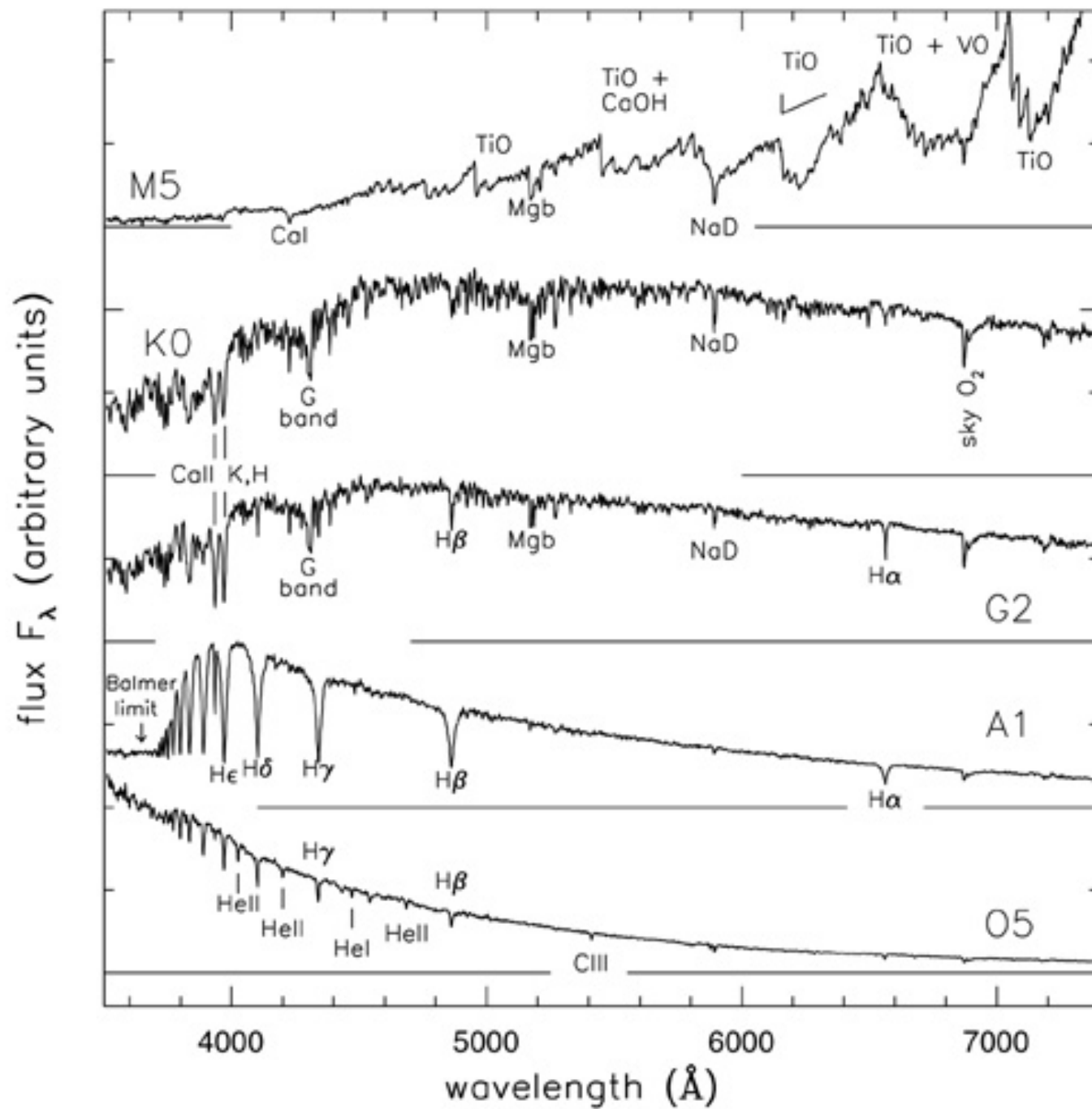


$$\frac{X^2}{1-X} = \frac{1}{n_H \Lambda^3} e^{-\chi/kT}$$

$$\chi = 13.6 \text{ eV}$$

$$X = \frac{n_e}{n_H}$$





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