Lab-03   **Bohr's Model**

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Name_____ **Sec/Group**_____ **Date**_____

TABLE-I

Emission Spectrum of Hydrogen

Line #	Measured Wavelength $\lambda_{meas.}$ (in *nm*)	Calculated Wavelength $\lambda_{calc.}$ (in *nm*)	Percent Difference $([\lambda_{meas} - \lambda_{calc}]/ \lambda_{calc})*100$
1			
2			
3			
4			
5			
6			

(Note: you may not be able to see six spectral lines. All rows need not be filled.)

Lab-03   **Bohr's Model**

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Name_____ **Sec/Group_____** **Date_____**

TABLE-II

Emission Spectrum of Helium

Line #	Measured Wavelength λ_{meas}. (in *nm*)	Calculated Wavelength λ_{calc}. (in *nm*)	Percent Difference $([\lambda_{meas} - \lambda_{calc}]/ \lambda_{calc})*100$
1			
2			
3			
4			
5			
6			

(Note: you may not be able to see six spectral lines. All rows need not be filled.
If you can not identify a line, you need not calculate 'Percent Difference' for that line)

    Write a brief explanation why the agreement between the theoretical prediction of
Bohr's model and the experimental observation is much better in the case of hydrogen
than for helium.